First-principles predictions of potential hydrogen storage materials:
Nanosized Ti(core)/Mg(shell) hydrides

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MgH₂ is one of the most promising hydrogen storage materials. However MgH₂ is thermodynamically too stable, leading to a too high desorption temperature of 300 °C at atmospheric pressure, which is a major impediment for practical applications. In this study, aiming to tune the thermodynamic stability of the MgH₂, nanosized two-dimensional Mg/Ti/Mg sandwich and three-dimensional Ti(core)/Mg(shell) hydrides have been investigated by using density functional theory calculations. For both structures, four types of hydrogen atoms can be distinguished: on the surface of the Mg (Hsurf), within the Mg (HMg), at the Mg/Ti interface (HMgTi), and within the Ti (HTi). For the dehydrogenation reaction, the hydrogen desorption from the hydride is in the order Hsurf, HMg, HMgTi, HTi. The desorption energy of Hsurf is unexpectedly high. As expected, due to the well-preserved fluorite structure of the partially hydrogenated hydride, the desorption energy of HMg is significantly lower than that of bulk rutile MgH₂. The further desorption of HMgTi and HTi becomes more difficult due to the strong Ti-H bonding. We propose that partial hydrogenation without adsorption of Hsurf and partial dehydrogenation without desorption of HMgTi and HTi would keep the fluorite symmetry with its favorable thermodynamics. The reversible hydrogen capacity (HMg) of the Mg/Ti/Mg sandwich structure is low, whereas the reversible hydrogen capacity of the Ti(core)/Mg(shell) is calculated to be reasonable high. Our results predicted Ti(core)/Mg(shell) structures as potential useful materials for hydrogen storage application.

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I. INTRODUCTION

MgH₂ is one of the most attractive hydrogen storage materials because it is inexpensive and light.1–3 However, the kinetics of the hydrogen uptake and release in Mg is poor. It is generally accepted that the diffusion limitations, high oxidation sensitivity of the Mg surface, and high activation barrier of hydrogen molecule dissociation lead to poor hydrogen kinetics in Mg. Another disadvantage of MgH₂ is too stable, leading to too high desorption temperatures (Td) at atmospheric pressure. For example, the desorption energy of MgH₂ is 75 kJ/mol H₂, i.e., 0.78 eV/H₂ (Td = 300 °C),4 while a desorption energy of 20–50 kJ/mol H₂, i.e., 0.21–0.52 eV/H₂ (Td = 20–100 °C) is desirable. Tuning kinetics of MgH₂ has been achieved by means of alloying Mg with transition metals (TMs), adding catalyst and mechanical ball milling.5–28 Among the Mg-TM-H hydrides, Mg-Ti-H is one of the most widely investigated systems because of its tremendous potential for improving the (de)hydrogenation kinetics.18–21 The fcc symmetry of the Mg-Ti-H system has been claimed to be responsible for the improved kinetics.19–21 The structural transformation from rutile to fluorite in Mg-Ti-H has been theoretically confirmed for both bulk and multilayered systems.31

Besides the improved kinetics, tuning of the thermodynamics of Mg-Ti-H has been reported for various morphologies, i.e., nanostructured particles,26 thin films,27 and multilayers.28 Nanostructured Mg-Ti-H by mixing MgH₂ and TiH₂ and ultrahigh energy high-pressure mechanical milling has been prepared by Lu et al.26 Their results show that the dehydrogenation energy of MgH₂-0.1TiH₂ is lower than that of MgH₂. For the rehydrogenation of the dehydrogenated alloys, fast hydrogen absorption was also observed at room temperature. Baldi et al.27,28 managed to tune the thermodynamics of hydrogen absorption in the Mg-Ti-Pd system by means of elastic clamping. Their results show that the thermodynamics of hydrogen absorption in Pd-capped Mg films are strongly dependent on the Mg thickness.27 Furthermore, Mg/Ti multilayers with various monolayer thicknesses between 0.5 and 20 nm were prepared. The layer thickness dependence of the hydrogenation properties was reported. Besides the elastic clamping effect of the Pd layer, the interface effect was also proposed to be responsible for the different thermodynamics.28

Despite all the experimental progress, the origin of the tuned thermodynamics of Mg-H bonding is still not well understood. Theoretical studies32–37 have reported that using nanoparticles instead of bulk or coarse particles of hydrides can alter the thermodynamics of hydrogen uptake and release. Because the energies of both metals and hydrides change going from bulk materials to nanosized particles, the reaction thermodynamics will be affected by the size of the grains. Moreover, our previous theoretical study has shown that the structural transformation of MgH₂ from rutile to fluorite induced by contacting with a TM plays an important role in tuning the thermodynamics in the MgH₂/TMH₂ multilayers.31 In MgH₂/TMH₂ multilayers, notable destabilization of Mg-H bonding could be observed with a high Ti:Mg ratio and very thin Mg layers (a few nanometers). This, however, leads to a significant reduction of the reversible hydrogen storage capacity due to the fact that parts of the hydrogen atoms are trapped in the TM. Stimulated by the promising results reported above, we consider combining the destabilized Mg-H bondings in the fluorite structure by alloying with Ti, and the surface effect in the nanosized materials. We propose that by extending the two-dimensional (2D) structure to three
dimensions (3D), the kinetics and thermodynamics can be improved and that a larger reversible hydrogen storage capacity can be expected. To verify this hypothesis, 2D Mg/Ti/Mg sandwich and 3D Ti(core)/Mg(shell) structures have been designed and studied by large-scale density functional theory (DFT) calculations. We intend to give a prediction of the thermodynamic properties in such nanostructured materials, thus helping to design unique hydrogen storage materials.

II. STRUCTURAL MODELS AND COMPUTATIONAL METHODS

Figure 1 shows the models for the Mg/Ti/Mg sandwich and Ti(core)/Mg(shell) structures. For simplicity, the 2D slab is cut from fluorite hydride with a (100) plane and the 3D particle is cut from fluorite hydride with a cubic symmetry. Four types of hydrogen atoms can be recognized in both structures: hydrogen on the surface of the Mg (Hsurf), within the Mg (HMg), at the Mg/Ti interface (HMGTi), and within the Ti (HTi). Three different sizes of the 2D (Fig. 2) and 3D (Fig. 3) structures are constructed by putting one, two, or three atomic layers of Mg on a fixed-sized Ti slab or Ti core.

All calculations were performed using DFT as implemented in the Vienna Ab Initio Simulation Package (VASP). The Kohn-Sham equations were solved using a basis of plane-wave functions with a plane-wave energy cutoff of 400 eV, and using pseudopotentials to describe the core electrons. The Perdew-Wang 1991 generalized gradient approximation was used for the electron-exchange correlation potential. For the 2D Mg/Ti/Mg sandwich structure (Fig. 2), (2 × 2) supercells were always used. A total of 7 × 7 × 1 k points were used to model the Brillouin zone. For the 3D Ti(core)/Mg(shell) structures, k-space integration is restricted to the Γ point only. Since periodic boundary conditions are involved, the size of the supercell containing the 2D nanoslab and 3D nanoparticle was chosen such that the distance between atoms of the periodic images exceeded 15 Å. This was sufficient to avoid interaction of the cluster with its images. Therefore, the calculations of Ti14Mg48, Ti14Mg158, and Ti14Mg350 (shown in Fig. 3) have been performed using a cubic supercell with sizes of 22 × 22 × 22, 30 × 30 × 30, and 38 × 38 × 38 Å3, respectively. For all structures, the lattice parameters, the volume, and the atom positions were allowed to relax.

III. RESULTS AND DISCUSSIONS

This section is divided into two parts. In the first part, the Mg/Ti/Mg sandwich structure is studied. The absorption sites and concentration of Hsurf and the absorption and desorption sequences of the four types of hydrogen atoms are investigated. After that, the size dependence of the dehydrogenation energies is studied. In the second part, extending the 2D structure to a 3D one, the Ti(core)/Mg(shell) structure is studied in a similar manner to the sandwich structure. We focus on the tuned thermodynamics, in particular, the relations between the thermodynamics and the structures and the Ti/Mg ratios. Furthermore, an oversaturation of the Ti(core)/Mg(shell) structure with extra Hsurf is studied by electronic and geometric structure calculations. Finally, the reversible capacities are determined and compared to the 2D structure.

A. Mg/Ti/Mg sandwich structure

1. Hsurf

Figure 4 displays the top view of the optimized Mg/Ti/Mg sandwich structure with five different distributions of Hsurf. After optimization, the hydrogen molecules in structure (a) and (b) have been “pushed” out from the Mg surfaces. The optimized Mg-H2 distances are 3.79 and 3.76 Å for (a) and (b), respectively. The desorption energies of these hydrogen molecules are nearly zero. This indicates that molecular hydrogen does not form chemical bonds with the Mg surface. For atomic hydrogen with a concentration of H/M = 2 shown in (c), half of the hydrogen atoms have associated to molecular hydrogen after optimization. These hydrogen molecules are far away from the Mg surface with a distance of ~3.50 Å. The remaining half of the hydrogen atoms are located closely to the Mg surface at short bridge sites. The optimized structures of (d) and (e) both demonstrate that short bridge sites with a concentration of H/M = 1 is the most stable structure. Altogether, considering two Mg surfaces per 2D slab, a hydrogen concentration of H/M = 2 for the whole system can be obtained. Therefore, the possibility of the oversaturated hydride can be excluded.

2. Hydrogen absorption and desorption sequences

To study the different thermodynamic properties for the four types of hydrogen atoms, hydrogen absorption and desorption energies in the Mg/Ti/Mg sandwich structure [Ti24Mg32Hx in Fig. 2(b)] have been calculated according to the following equations,

\[ \Delta E_{\text{Habsorption}} = E_{\text{Mg,Ti1-H}_x} - E_{\text{Mg,Ti1-H}_x} - \frac{z}{2} E_{\text{H}_2}, \]  

where \[ E_{\text{Mg,Ti1-H}_x}, \] \[ E_{\text{Mg,Ti1-H}_x}, \] \[ E_{\text{Mg,Ti1-H}_x}, \] and \[ E_{\text{Mg,Ti1-H}_x}, \] are the total energy of Mg, Ti1-Hx, and Mg, Ti1-Hx, normalized to the number of metal atoms.

The hydrogen absorption and desorption sequences are determined by comparing the hydrogen absorption and desorption energies at different hydrogen locations. The calculated energies are summarized in Table I. For instance,
Ti_{24}Mg_{32}H_{112−16} denotes that 16 hydrogen atoms are removed from each type of hydrogen location in the Ti_{24}Mg_{32}H_{112}. Four types of hydrogen are considered and the desorption energy of H_{surf} is the lowest, therefore H_{surf} will desorb first. After the desorption of H_{surf}, Ti_{24}Mg_{32}H_{96−32} denotes desorption of 32 hydrogen atoms of H_{Mg}, H_{MgTi}, or H_{Ti} from Ti_{24}Mg_{32}H_{96}. The desorption energy of H_{Mg} is the lowest, therefore it will desorb secondly. When there are only H_{MgTi} and H_{Ti} left, H_{MgTi} will desorb before H_{Ti}. Therefore, for the dehydrogenation reaction, the desorption of hydrogen follows the order of H_{surf}, H_{Mg}, H_{MgTi}, and H_{Ti}. For the stepwise dehydrogenation energies are 1.15 eV/H_{2} for H_{surf}, 0.38 eV/H_{2} for H_{Mg}, 1.30 eV/H_{2} for H_{MgTi}, and 1.10 eV/H_{2} for H_{Ti}. The desorption energy of H_{surf} is surprisingly high. They are many higher than the hydrogen desorption energy in bulk MgH_2 (0.65 eV/H_2). After the desorption of H_{surf}, the desorption of H_{Mg} becomes much easier with very low desorption energy. As expected, after the desorption of H_{surf} and H_{Mg}, the desorption of H_{MgTi} and H_{Ti} becomes more difficult due to the introduction of strong Ti-H bonding. The absorption sequence was determined in a similar manner: The location with the lowest absorption energy (H_{Ti}) absorbs hydrogen first; after the absorption of the H_{Ti}, the absorption of the H_{MgTi} is the most favorable in the remaining three. The absorption sequence is calculated to be the opposite to the desorption sequence: H_{Ti}, H_{MgTi}, H_{Mg}, H_{surf}.

### 3. The size dependence of the dehydrogenation energies

The different thermodynamic properties of the four types of hydrogen species predicted above indicates the potential to tune the thermodynamics of dehydrogenation reaction of the 2D Mg/Ti/Mg sandwich structure. In particular, the populations of different types of hydrogen depend highly on the size and composition of the structure. Average and stepwise dehydrogenation energies of the Ti_{24}Mg_{16}H_{80} [Fig. 2(a)] and Ti_{24}Mg_{48}H_{144} [Fig. 2(c)] have been calculated and compared to those of the Ti_{24}Mg_{32}H_{112} [Fig. 2(b)] studied in Sec. III A 2. The calculated hydrogen desorption sequence of H_{surf}, H_{Mg}, H_{MgTi}, and H_{Ti} in Sec. III A 2 were applied.

As one can see in Table II, the dehydrogenation energies of all the three structures with the desorption of H_{surf} are surprisingly high. The reason will be explained further with the structural analyses in Sec. III B 2. Similar to that of Ti_{24}Mg_{32}H_{112}, the desorption of H_{Mg} for Ti_{24}Mg_{32}H_{112}

<table>
<thead>
<tr>
<th>Desorption</th>
<th>H_{surf}</th>
<th>H_{Mg}</th>
<th>H_{MgTi}</th>
<th>H_{Ti}</th>
<th>Absorption</th>
<th>H_{surf}</th>
<th>H_{Mg}</th>
<th>H_{MgTi}</th>
<th>H_{Ti}</th>
</tr>
</thead>
<tbody>
<tr>
<td>Ti_{24}Mg_{32}H_{112−16}</td>
<td>1.15</td>
<td>1.56</td>
<td>1.45</td>
<td>1.60</td>
<td>Ti_{24}Mg_{32}H_{16−16}</td>
<td>0.09</td>
<td>0.48</td>
<td>−1.10</td>
<td>−1.21</td>
</tr>
<tr>
<td>Ti_{24}Mg_{32}H_{96−32}</td>
<td>—</td>
<td>0.38</td>
<td>1.40</td>
<td>1.49</td>
<td>Ti_{24}Mg_{32}H_{22−16}</td>
<td>−0.43</td>
<td>−0.19</td>
<td>−1.39</td>
<td>—</td>
</tr>
<tr>
<td>Ti_{24}Mg_{32}H_{44−32}</td>
<td>—</td>
<td>—</td>
<td>1.30</td>
<td>1.38</td>
<td>Ti_{24}Mg_{32}H_{64−16}</td>
<td>−0.31</td>
<td>−0.36</td>
<td>—</td>
<td>—</td>
</tr>
<tr>
<td>Ti_{24}Mg_{32}H_{32−32}</td>
<td>—</td>
<td>—</td>
<td>—</td>
<td>1.10</td>
<td>Ti_{24}Mg_{32}H_{66−16}</td>
<td>−1.11</td>
<td>—</td>
<td>—</td>
<td>—</td>
</tr>
</tbody>
</table>
FIG. 3. (Color online) Optimized 3D Ti(core)/Mg(shell) alloys and their fully and partially hydrogenated hydrides involved in the (de)hydrogenation processes. For each Ti/Mg composite, the structures from the left-hand side to the right-hand side indicate the dehydrogenation process.

(0.36 eV/H₂) is very low and after the desorption of Hsurf and HMg. Because of the strong Ti-H bonding, the desorption energies of HMgTi and HTi are always larger than the rutile MgH₂ (0.65 eV/H₂). To save computational time, we combined the desorption of HMgTi with HTi and only the average desorption of the two was calculated.

Comparing the three hydrides, one can notice the particle size dependence of the hydrogen thermodynamics. On the one hand, the average dehydrogenation energies of all the hydrogen atoms are decreasing with increasing size of the Mg layer, i.e., from 1.22, 0.97, and 0.91 eV/H₂ for Ti₂₄Mg₁₆H₈₀, Ti₂₄Mg₃₂H₁₁₂, and Ti₂₄Mg₄₈H₁₄₄, respectively. This can be explained as follows: The larger the population of the Mg-related sites, and the smaller the average hydrogen desorption energy. On the other hand, the desorption energy of Hsurf increases with increasing Mg layer thickness, from 0.95 eV/H₂ for Ti₂₄Mg₁₆H₈₀, to 1.15 eV/H₂ for Ti₂₄Mg₃₂H₁₁₂, to 1.42 eV/H₂ for Ti₂₄Mg₄₈H₁₄₄. The reason for the unexpectedly high desorption energies of Hsurf and the size dependence will be explained together with the 3D Ti(core)/Mg(shell) structure in Sec. III B 2.

It is worth noting that, for Ti₂₄Mg₃₂H₁₁₂ and Ti₂₄Mg₄₈H₁₄₄, the desorption energies of HMg (0.38 and 0.36 eV/H₂) are much lower than that of bulk rutile MgH₂ (0.65 eV/H₂). The desorption energy of HMg is low because the HMg always occupy the tetrahedral sites in the partially dehydrogenated fluorite type hydrides [see Figs. 2(b)(ii) and 2(c)(ii)]. This is favorable for hydrogen storage application. However, the population of these destabilized hydrogen atoms (HMg) is low: 1.7 wt % for Ti₂₄Mg₃₂H₁₁₂ and 2.7 wt % for Ti₂₄Mg₄₈H₁₄₄. To obtain a larger population of HMg, we extend the 2D Mg/Ti/Mg sandwich structure to a 3D Ti(core)/Mg(shell) structure in the next section.

| TABLE II. Stepwise and average dehydrogenation energies (in eV/H₂) of three Mg/Ti/Mg sandwich hydrides with different particle sizes and compositions. The hydrogenation energies are not shown as they are identical to the dehydrogenation energy apart from the sign change. |
|---|---|---|---|---|
| Fig. 2 Mg/Ti/Mg sandwich | Hsurf | H₂|Mg | Hsurf+H₂|Mg + HTi | Average |
| (a) Ti₂₄Mg₁₆H₈₀ | 0.95 | — | 1.29 | 1.22 |
| (b) Ti₂₄Mg₃₂H₁₁₂ | 1.15 | 0.38 | 1.31 | 0.97 |
| (c) Ti₂₄Mg₄₈H₁₄₄ | 1.42 | 0.36 | 1.33 | 0.91 |

B. Ti(core)/Mg(shell) structure

Because the same hydrogen locations can be found in the 2D and 3D structures, we assume that the adsorption sites of Hsurf, and absorption and desorption sequences found in the 2D structure, also apply for the 3D Ti(core)/Mg(shell) structure. However, it should be emphasized at this point that in the 2D structure, the Mg surfaces are infinitely extended two dimensionally. When hydrogen adsorbs on the Mg surface with a ratio of H/M = 1, an overall ratio of H/M = 2 can always
be obtained. However, for the Ti(core)/Mg(shell) structure slightly different situations were observed. During optimization, the hydrogen atoms adsorbed on the crystallographic facets were indeed stabilized at short bridge (sb) sites, whereas hydrogen atoms on the edges migrated from the sb to the long bridge (lb) sites. The number of sb and lb sites depend highly on the configuration of the core-shell alloy. For instance, 36 (24sb + 12lb), 144 (108sb + 36lb), and 228 (192sb + 36lb) sites are available for hydrogen atoms on the surfaces of Ti14Mg48, Ti14Mg158, and Ti14Mg350 (see Fig. 3), respectively. Together with H_{Mg}, H_{MgTi}, and H_{Ti}, the fully hydrogenated hydrides are Ti14Mg48H_{124}, Ti14Mg158H_{360}, and Ti14Mg350H_{740}, i.e., H/M = 2 or slightly larger than 2. Four single-point calculations with a concentration of Ti14Mg158H_{352} (where eight hydrogen atoms of each type were taken out from Ti14Mg158H_{360}) were calculated. The results show the same desorption sequence with what was predicted for the 2D sandwich structure. Therefore, we are confident that the 3D (de)hydrogenation process is similar to that of the 2D structure.

1. The stepwise dehydrogenation energies

The average and stepwise dehydrogenation energies were calculated in the same manner as for the 2D structure. As one can see in Table III, the trends found for the energies in the 2D structure are reproduced: The desorption energies of H_{surf} are very high (except that of Ti14Mg48H_{124}); after the desorption of H_{surf}, the desorption energies of H_{Mg} become lower than that of bulk MgH_2; the desorption of H_{MgTi} + H_{Ti} become more difficult due to the strong Ti-H bondings. It is worth noting that for Ti14Mg158H_{360} and Ti14Mg350H_{740}, the desorption energies of H_{Mg} (0.39 and 0.27 eV/H_2) are much lower than that of the rutile MgH_2 (0.65 eV/H_2). They are within the desired thermodynamics range and even lower than the lowest H_{Mg} desorption energy of 0.44 eV/H_2 predicted for the MgH_2/TiH_2 multilayers. The desorption energy of H_{Mg} is low because the H_{Mg} atoms always occupy the tetrahedral sites in the partially dehydrogenated hydrides [shown in Figs. 3(b)(ii), and 3(c)(ii)]. This is favorable for hydrogen storage applications.

Comparing the three hydrides, one can again notice the same size dependence of the energies to that of the 2D structure. On the one hand, the average dehydrogenation energies of all hydrogen atoms are decreasing with increasing size of the Mg shell, i.e., 0.80, 0.68, and 0.65 eV/H_2 for Ti14Mg48H_{124}, Ti14Mg158H_{360}, and Ti14Mg350H_{740}, respectively. On the other hand, the desorption energy of H_{surf} increases with increasing size of the Mg shell, from 0.59 eV/H_2 for Ti14Mg48H_{124}, to 1.11 eV/H_2 for Ti14Mg158H_{360}, to 1.42 eV/H_2 for Ti14Mg350H_{740}. The reason for the unexpectedly high values and the size dependence of the desorption energies of H_{surf} will be explained by comparing the structures of the fully and partially hydrogenated hydrides in Sec. III B 2.

2. Structures

To explain the trend found in the dehydrogenation energies of H_{surf} in both 2D and 3D structures, Figs. 2 and 3 should be referred. As one can see, for the fully hydrogenated hydrides, the H_{surf} atoms do not occupy the perfect short bridge site. The hydrogen atoms at the subsurface are also dislocated from perfect tetrahedral sites. These distortions are similar with the deformations which were found in the Mg/Ti multilayers, where with the increasing Mg thickness the fluorite structure tends to change into rutile. For instance, approximately half of the H_{Mg} atoms are closely bonded to H_{surf}. These atoms are ~0.2–0.4 Å closer to the surface Mg than the other half, which are less affected by H_{surf}. This effect extends from one layer to another through the H-H bonding, i.e., from the subsurface of the Mg layer to the Mg/Ti interface. It becomes minor for H_{MgTi} and H_{Ti} atoms as they are stabilized by the strong Ti-H

### Table III. Stepwise and average dehydrogenation energies (in eV/H_2) of the Ti(core)/Mg(shell) structure.

<table>
<thead>
<tr>
<th>Fig. 3 Ti(core)/Mg(shell)</th>
<th>H_{surf}</th>
<th>H_{Mg}</th>
<th>H_{MgTi} + H_{Ti}</th>
<th>Average</th>
</tr>
</thead>
<tbody>
<tr>
<td>(a) Ti14Mg48H_{124}</td>
<td>0.59</td>
<td>—</td>
<td>0.95</td>
<td>0.80</td>
</tr>
<tr>
<td>(b) Ti14Mg158H_{360}</td>
<td>1.11</td>
<td>0.39</td>
<td>0.90</td>
<td>0.68</td>
</tr>
<tr>
<td>(c) Ti14Mg350H_{740}</td>
<td>1.42</td>
<td>0.27</td>
<td>0.80</td>
<td>0.65</td>
</tr>
</tbody>
</table>
bonding. Therefore, the structural deformations depend on the size of the Mg layer. The thicker the layer, the more HMg atoms displaced by Hsurf. The structural deformation is direct evidence that HMg atoms are stabilized upon forming the H-H bonding by the absorption of Hsurf. This also explains why HMg becomes unstable upon the desorption of Hsurf in Sec. III A 2.

However, after the desorption of Hsurf from the Mg surface, HMg atoms will relocate at the tetrahedral sites (see Figs. 2 and 3). A well-defined fcc structure is formed again. The fcc symmetry can be preserved in the continuing dehydrogenation process. The desorption energies of Hsurf were calculated by comparing the energies of the fully hydrogenated and partially hydrogenated hydrides. In this case, they consist of the energy cost of the breaking of the Mg-Hsurf bond and the energy difference of the Mg-HMg bonding between the fully and partially hydrogenated hydrides. The latter closely relates to the population of HMg. Thus the thicker the Mg layer, the more HMg is stabilized in the fully hydrogenated hydrides, and the larger the Hsurf desorption energy.

3. Oversaturation of the Ti(core)/Mg(shell) structure

The hydrogen storage capacities in nanosized materials can vary from their corresponding bulk materials. For instance, Kobayashi et al. reported that the capacity of the Pd-Pt nanoparticles with a Pt content of 8–21 at. % was higher than that of pure Pd nanoparticles. The amount of hydrogen absorption is controllable by altering the alloy compositions of Pd and Pt. Wagemans et al. predicted that small MgH2 particles (smaller than Mg13H30) have the potential to take up 10%–15% extra hydrogen. They proposed these extra hydrogen atoms are less strongly adsorbed onto the surface of the cluster than the hydrogen absorbed within the cluster. The above results suggest oversaturation of the surface metal.

To verify the potential extra hydrogen in the present nanosized Ti(core)/Mg(shell) alloys, an oversaturation of the Ti14Mg158H512 alloy has been studied. Figure 5(a) shows the initial structure of Ti14Mg158H512 (H/M = 2.98), i.e., the fluorite structure with hydrogen terminations. As shown in Fig. 5(b), during the geometric optimization the hydrogen atoms on the Mg surface (Hsurf) did not dramatically distort the structure of the alloy, but formed different types of hydrogen species. They can be classified as the hydrogen dimer (H-H), hydrogen trimer (H-H-H), hydrogen atoms on facets (Hface), and hydrogen atoms at vertices (Hver and Hver2). Except for the hydrogen atoms on the Mg shell surface, the other ones occupy different tetrahedral sites within the Mg shell (HMg), at the Ti/Mg interface (HMgTi), and within the Ti core (HTi).

To study the hydrogen bonding properties, the partial density of states (PDOS) of all types of hydrogen atoms in Ti14Mg158H512 have been calculated [see Fig. 5(c)]. In the PDOS plot, a sharp peak ∼−7.5 eV can be found for the hydrogen dimer. The extremely localized density clearly indicates that no chemical bonds are formed between hydrogen molecules and the alloy. In fact, the hydrogen molecules are “pushed” out of the Mg surface with distances of 2.42 Å (facets) and 2.36 Å (edges), respectively. At vertices, each Mg atom is surrounded by one Hver and three Hverb. The Hver is very close to the Mg with a distance of 1.71 Å, whereas the other three Hverb are further with a distance of 1.90 Å.
means that only H_Mg will be reversible in the hydrogen cycles. Such reversible hydrogen capacities are estimated to be reasonable high: 3.3 and 4.8 wt % for Ti_{14}Mg_{158}H_{360} and Ti_{14}Mg_{350}H_{450}, respectively. They are much higher than the reversible capacity of 1.4 wt % in the Mg/Ti multilayers with a H_Mg desorption energy of 0.44 eV/H_2,\textsuperscript{31} and also higher than the ones predicted in the Mg/Ti/Mg sandwich structures in Sec. III A 3 (1.7 and 2.7 wt %).

IV. CONCLUSIONS

Hydrogen storage properties of nanosized 2D Mg/Ti/Mg sandwich and 3D Ti(core)/Mg(shell) hydrides have been studied using large-scale DFT calculations. For both structures, four types of hydrogen atoms can be distinguished in the Ti(core)/Mg(shell) hydrides: H_{surf}, H_Mg, H_MgTi, and H_Ti. For the hydrogenation reaction, the hydrogen absorption in the alloys is in the order of H_Ti, H_MgTi, H_Mg, and H_{surf}. For the dehydrogenation reaction, the hydrogen desorption is in the reverse order. The desorption energies of H_{surf} are calculated to be high. The reason is the structural transformation from fluoride to a deformed fluoride structure upon the adsorption of H_{surf}. In other words, the absorption of H_{surf} induces dislocation of H_Mg and stabilizes the hydrides. This stabilization effect increases with increasing thickness of the Mg layer. Due to the well-preserved fluorite structure of the partially hydrogenated hydrides, the desorption energies of H_Mg are significantly lower than that of bulk MgH_2. This is favorable for hydrogen storage applications. But the further desorption of H_MgTi and H_Ti is more difficult than that of bulk MgH_2 due to the strong Ti-H bonding. Furthermore, oversaturation of the 3D Ti(core)/Mg(shell) structures is studied. Both the geometric and electronic structure analyses show that the oversaturation of the structure is not stable.

Based on the observed thermodynamic properties, partial hydrogenation without the adsorption of H_{surf} and partial dehydrogenation without the desorption of H_MgTi and H_MgTi is proposed to preserve the fluorite structure of the hydride with desired thermodynamics. However, this leads to a reduction of the reversible hydrogen capacity. The reversible hydrogen capacity (H_{Mg}) of the Mg/Ti/Mg sandwich structures are low: 1.7 wt % for Ti_{24}Mg_{32}H_{112} and 2.7 wt % for Ti_{24}Mg_{48}H_{144}. Those of the Ti(core)/Mg(shell) structures are estimated to be reasonable high: 3.3 and 4.8 wt % for Ti_{14}Mg_{158}H_{360} and Ti_{14}Mg_{350}H_{450}, respectively. Our results indicate that the nanosized Ti(core)/Mg(shell) hydrides are very promising hydrogen storage materials with favorable thermodynamics and reasonable hydrogen capacity.